

Page 5799. The authors wish to report further: "On repeating the experimental work, R. J. Steltenkamp obtained somewhat different results. The principal ethoxyacetylene adduct (b.p.  $107^\circ$  (3 mm.),  $n_D^{25}$  1.5500; hydrolysis gave *p*-toluenethiol and no dinitrophenylhydrazone of *p*-tolylmercaptoacetaldehyde; aniline gave *N,N'*-diphenylacetamide, m.p.  $131$ – $132.5^\circ$ ) is different from  $C_7H_7SCH=CHOEt$  prepared independently from *cis*- $C_7H_7SCH=CHCl$  (fair conversion), from  $C_7H_7SCH_2CH(OEt)_2$  (poor conversion) and *via* the very clean free-radical addition of  $C_7H_7SH$  to ethoxyacetylene (b.p.  $102^\circ$  (0.55 mm.),  $n_D^{25}$  1.5640; yielded the dinitrophenylhydrazone of *p*- $C_7H_7SCH_2CHO$ ). Hence the principal initial ethoxyacetylene adduct under Goldhamer's conditions is  $CH_2=C(OEt)SC_7H_7$  and not *cis*- $C_7H_7SCH=CHOEt$ . This initial adduct shows essentially one peak on a gas chromatogram (one of the two minor peaks comes from  $CH_2COSC_7H_7$ ), but on standing this peak becomes minor and several others appear, the major one of which derives from  $C_7H_7SCH=CHOEt$ , and another from  $C_7H_7SH$ . Presumably the initial 1,1-adduct reverts to starting materials, which recombine in a free radical process."—WILLIAM E. TRUCE.

**Henry Feuer and Christos Savides.** Reactions of  $\alpha,\alpha'$ -Dibromo- $\alpha,\alpha'$ -dinitrodinitriles.

Page 5831. In col. 1, line 1, for "259" read "299."—HENRY FEUER.

**C. J. Morris, J. F. Thompson, S. Asen and F. Irreverre.** Isolation of a New Acidic Aromatic Amino Acid (*m*-Carboxy- $\alpha$ -phenylglycine) from Iris Bulb.

Page 6069. This article should have been recorded as a joint contribution by C. J. Morris and J. F. Thompson of the U. S. Plant, Soil and Nutrition Laboratory of the U. S. Department of Agriculture at Ithaca, N. Y., S. Asen of the Crops Research Division of the U. S. Department of Agriculture at Beltsville, Md., and F. Irreverre of The National Institute of Arthritis and Metabolic Diseases, National Institutes of Health, Bethesda, Md.

**Gary Felsenfeld and Morton P. Printz.** Specific Reactions of Hydrogen Peroxide with the Active Site of Hemocyanin. The Formation of "Methemocyanin."

Page 6261. In col. 1, line 16, the mobilities are  $-3.32$  and  $-3.44$   $cm.^2$   $sec.^{-1}$   $volt^{-1} \times 10^{-5}$ , measured at  $4^\circ$ .

1960, VOL. 82

**Stanley J. Cristol and Robert S. Bly, Jr.** Mechanisms of Elimination Reactions. XXI. The Alkaline Dehydrochlorination of *meso*- and *dl*-Stilbene Dichlorides.

Page 144. In col. 2, text line 7 from the end, for "destabilizes" read "stabilizes."—STANLEY J. CRISTOL.

**H. James Harwood and Daniel W. Grisley, Jr.** The Unexpected Course of Several Arbuzov-Michaelis Reactions; an Example of the Nucleophilicity of the Phosphoryl Group.

Page 423. In col. 2, line 17, after "tion" add "of this latter ester with alkyl halide. This latter reaction."—H. JAMES HARWOOD.

**R. J. Alexander, N. C. Baenziger, C. Carpenter and J. R. Doyle.** Metal-Olefin Compounds. I. The Preparation and Molecular Structure of Some Metal-Olefin Compounds Containing Norbornadiene (Bicyclo[2.2.1]hepta-2,5-diene).

Page 535. The name of the first author should be "R. J. Alexander."

Page 537. In col. 1, line 43, for "15" read "16."—J. R. DOYLE.

**E. L. Muettterties.** Stereochemistry of Complexes Based on Metal Tetrafluorides.

Page 1085. Col. 1, Table I, line 12, for " $MoF_4 \cdot C_2H_5N$ " read " $MoF_4 \cdot 2C_2H_5N$ ." Line 13, for " $MoF_4 \cdot N(CH_3)_2$ " read " $MoF_4 \cdot 2N(CH_3)_2$ ."—EARL L. MUETTERTIES.

**Richard H. Wiley, E. Imoto, R. P. Houghton and P. Veeravagu.** Synthesis and Characterization of the Geometric and Structural Isomers of 3,7-Dimethyl-2,4,6-octatrienoic Acid. I.

Page 1416. In col. 1, paragraph on Compound (VIII), lines 11–13 from the end, revise text with newer information:

"The ultraviolet absorption maximum for 3,7-dimethyl-2-*cis*:4-*trans*:6-octatrienoic acid (VIII), m.p.  $161^\circ$ , in methanolic 0.0002 *N* sulfuric acid is at  $311 m\mu$  ( $\log \epsilon$  4.54). The value  $300 m\mu$  for the maximum as printed was taken in methanol and for presumably partially ionized acid."—RICHARD H. WILEY.

**James E. Griffiths and Anton B. Burg.** The Phosphinous Acid  $(CF_3)_2POH$  and the Diphosphoxane  $(CF_3)_2POP(CF_3)_2$ .

Page 1507. In col. 2, line 10, for " $R_2POH$ " read " $R_2PO$ ."

Page 1508. In col. 2, line 18, for "too" read "to be."

**Costas H. Issidorides, Mary Fieser and Louis F. Fieser.** Selenium Dioxide Oxidation of Methyl  $\Delta^3$ -Cholenate.

Page 2002. In col. 2, lines 4 and 5, transpose " $155^\circ$ " and " $170^\circ$ ."—C. H. ISSIDORIDES.

**Teruo Matsuura and H. J. Cahnmann.** Model Reactions for the Biosynthesis of Thyroxine. III. The Synthesis of Hindered Quinol Ethers and their Conversion to Hindered Analogs of Thyroxine.

Page 2058. In col. 1, line 15, after "drogen" add the wording "or bromine are white crystalline substances. The analog in which  $R_1$  is an acetic acid side chain and  $R_2$  hydrogen."—TERUO MATSUURA AND H. J. CAHNMANN.

**Henry H. Bauer, David L. Smith and Philip J. Elving.** The Faradaic Admittance of Electrochemical Processes. II. Experimental Test for the Theoretical Equations.

Page 2095. The third and fourth lines of the legend to Figure 2 should read " $(1/k)\sqrt{\omega D/2}$  with  $k = 0.1$  and  $0.21$ , respectively; III, IV, V, theoretical lines for  $\cot \phi = 1 - (1/k)\sqrt{\omega D/2}$  with  $k =$ ."

Page 2097, the latter part of footnote *d* to Table II (third line) should read, "in result 13,  $\cot \phi = 1 - \frac{1}{k}\sqrt{\frac{\omega D}{2}}$ ."

Page 2098. In col. 1, line 4 from the end, for " $\phi$ " read " $\theta$ ."—PHILIP J. ELVING.

**John S. Meek, Donovan R. Wilgus and John R. Dann.** Diels-Alder Reactions of 9-Substituted Anthracenes. IV. 9-Nitroanthracene and 9-Anthramide.

Page 2568. In col. 2, line 24, for "9,11-" read "9,12-" and in line 26 for "half amide" read "imide."—JOHN S. MEEK.

**M. L. Wolfrom and Bienvenido O. Juliano.** Chondroitin Sulfate Modifications. II. Sulfated and *N*-Deacetylated Preparations.

Page 2588. The Roman numeral in the title should be III and footnote (1) should read: Part II, *J. Org. Chem.*, **25**, 308 (1960).—M. L. WOLFROM.

**David H. Geske and August H. Maki.** Electrochemical Generation of Free Radicals and their Study by Electron Spin Resonance Spectroscopy; the Nitrobenzene Anion Radical.

Page 2674. In col. 1, the line above the heading "Discussion," for "2.0032" read "2.0044."—DAVID H. GESKE.

**William A. Pryor.** The Kinetics of Oxidation by Aqueous Polysulfide Solutions. III. The Effect of Substituents on the Oxidation of Arylmethyl Compounds.

Page 2717. In Table V, line 3 of the footnotes, insert *d* before "See note *d* . . ." and for "*d* See ref. 24;" read "*e* See ref. 24;."—WM. A. PRYOR.

**Donald S. Matteson, J. J. Drysdale and W. H. Sharkey.** Thermal Rearrangement of  $\delta$ -Methylenecyclo[2.2.1]hept-2-ene.

Page 2854. In col. 2, line 1, for "consistent with a planar triene" read "inconsistent with a planar conjugated triene."—W. H. SHARKEY.

**Harold E. Zaugg, Bruce W. Horrom and Sandra Borgwardt.** Specific Solvent Effects in the Alkylation of Enolate Anions. I. The Alkylation of Sodiomalonic Esters with Alkyl Halides.